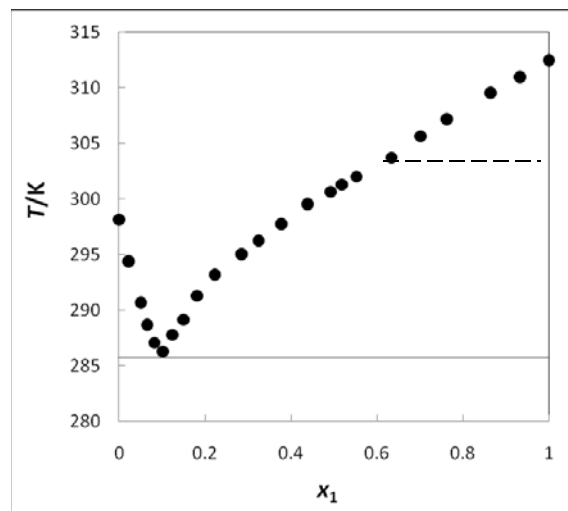


Example Table:  
SLE Data (with eutectic formation)

**Table 2. Experimental (Solid + Liquid) Equilibrium Temperatures  $T$  as a Function of Mole Fraction  $x$  for {1,4,7,10,13,16-Hexaoxacyclooctadecane (1) + tert-Butyl Alcohol (2)} at Pressure  $p = 0.1$  MPa <sup>a</sup>**

$x_1$	$T/K$	Solid phase
0.0000	298.15	tert-Butyl Alcohol(cr)
0.0225	294.40	tert-Butyl Alcohol(cr)
0.0508	290.70	tert-Butyl Alcohol(cr)
0.0658	288.70	tert-Butyl Alcohol(cr)
0.0826	287.10	tert-Butyl Alcohol(cr)
0.1022	286.30	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(II)}
0.1241	287.80	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(II)}
0.1498	289.15	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(II)}
0.1810	291.30	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(II)}
0.2234	293.20	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(II)}
0.2850	295.05	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(II)}
0.3249	296.25	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(II)}
0.3769	297.75	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(II)}
0.4389	299.55	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(II)}
0.4920	300.65	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(II)}
0.5183	301.30	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(II)}
0.5523	302.00	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(II)}
0.6336	303.72	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(I)}
0.7005	305.65	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(I)}
0.7620	307.20	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(I)}
0.8635	309.55	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(I)}
0.9320	310.95	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(I)}
1.0000	312.45	1,4,7,10,13,16-Hexaoxacyclooctadecane {cr(I)}



The figure is shown to illustrate the experimental data. This is **not** a standard format for this journal.

<sup>a</sup> Standard uncertainties are  $u(T) = 0.1$  K,  $u(x) = 0.0005$ ,  $u(p) = 5$  kPa.

**NOTE:** In the figure, the horizontal line near 285 K indicates the eutectic temperature, and the dashed line near 303 K indicates the cr(II)-to-cr(I) phase transition temperature for 1,4,7,10,13,16-hexaoxacyclooctadecane.