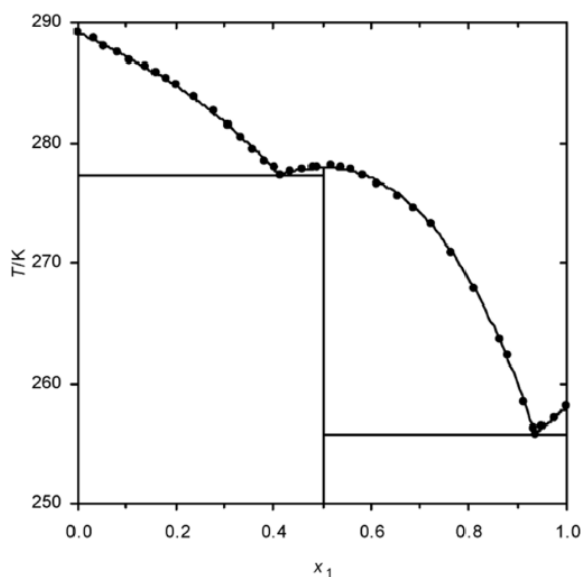# Example Table:
SLE Data (with compound formation)

TABLE 2
Experimental (solid + liquid) equilibrium temperatures T and liquid mole fractions x for the system octan-1-ol (1) + decylamine (2) at pressure p = 0.1 MPa.[a,b]

x_1	T/K	Solid phase	x_1	T/K	Solid phase
0.0000	289.16	Decylamine(cr)	0.4901	277.96	AB(cr)
0.0310	288.63	Decylamine(cr)	0.5167	278.00	AB(cr)
0.0556	288.06	Decylamine(cr)	0.5382	277.95	AB(cr)
0.0811	287.44	Decylamine(cr)	0.5603	277.70	AB(cr)
0.1087	286.87	Decylamine(cr)	0.5850	277.25	AB(cr)
0.1389	286.30	Decylamine(cr)	0.6122	276.60	AB(cr)
0.1590	285.76	Decylamine(cr)	0.6529	275.61	AB(cr)
0.1816	285.27	Decylamine(cr)	0.6883	274.57	AB(cr)
0.2006	284.76	Decylamine(cr)	0.7232	273.17	AB(cr)
0.2375	283.81	Decylamine(cr)	0.7648	270.80	AB(cr)
0.2779	282.58	Decylamine(cr)	0.8124	267.85	AB(cr)
0.3080	281.41	Decylamine(cr)	0.8652	263.60	AB(cr)
0.3343	280.42	Decylamine(cr)	0.8790	262.35	AB(cr)
0.3587	279.36	Decylamine(cr)	0.9149	258.45	AB(cr)
0.3821	278.45	Decylamine(cr)	0.9333	256.16	AB(cr)
0.4049	277.89	Decylamine(cr)	0.9526	256.35	Octan-1-ol(cr)
0.4345	277.60	AB(cr)	0.9753	257.14	Octan-1-ol(cr)
0.4623	277.75	AB(cr)	1.0000	258.03	Octan-1-ol(cr)
0.4832	277.83	AB(cr)			

[a] Standard uncertainties u are $u(T)$ = 0.1 K, $u(x)$ = 0.0005, and $u(p)$ = 5 kPa

[b] AB(cr) represents the crystal of the compound formed for mole fraction 0.5 of component 1.

The figure is shown to illustrate the experimental data. This is *not* a standard format for this journal.