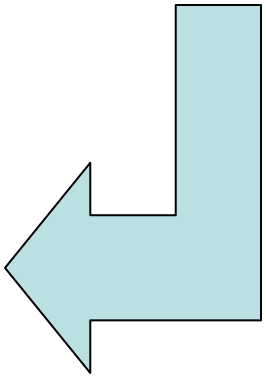
Table 1. Vapor Pressure of Barbituric Acid by Torsion Measurements

cell a								cell b1		cell b2	
run 1		run 2		run 3		run 4		run 1		run 1	
T/K	− log(*p*/kPa)	*T*/K	− log(*p*/kPa)	*T*/K	− log(*p*/kPa)	*T*/K	− log(p/kPa)	*T*/K	− log(*p*/kPa)	*T*/K	− log(*p*/kPa)
397	3.66	393	3.96	392	3.96	400	3.66	393	3.96	462	1.68
405	3.36	396	3.66	405	3.36	406	3.48	401	3.48	469	1.53
410	3.18	400	3.66	409	3.26	410	3.36	410	3.26	472	1.40
410	3.26	405	3.48	415	3.12	415	3.18	414	3.12	477	1.27
418	2.96	409	3.26	420	2.96	426	2.81	419	2.96	481	1.16
419	2.92	413	3.12	425	2.76	430	2.64	424	2.76	485	1.09
423	2.81	420	2.88	432	2.51	435	2.48	429	2.60	490	0.95
431	2.53	424	2.78	438	2.32	440	2.33	434	2.43	493	0.89
432	2.48	430	2.55	444	2.14			440	2.27		
436	2.34	436	2.39	449	1.99			445	2.12		
443	2.13	441	2.22					454	1.85		
453	1.84	447	2.04								
462	1.67	452	1.99								
		457	1.74								



Data Summary entry for this table

Table #: 1
System type (Pure, Binary, Ternary, Reaction): Pure
Chemical System(s): barbituric acid
Property: sublimation pressure
Experimental Method: torsion
Combined Expanded Uncertainty (*k* = 2) for the Property:
$2\sigma(p) = 0.5\ p$
State Variables and Constraints:
temperature T
Standard Uncertainty (*k* = 1) for Variables and Constraints:
$\sigma(T) = 2$ K